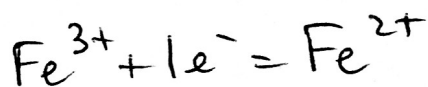
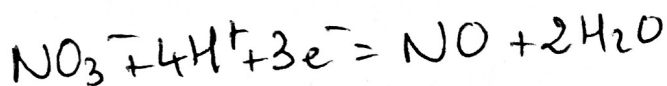
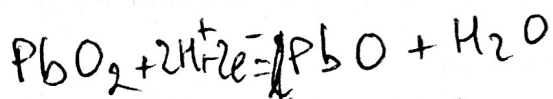
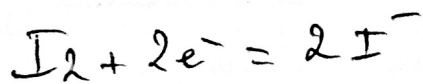
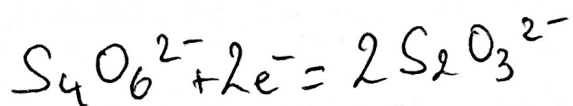
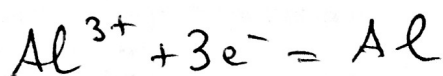
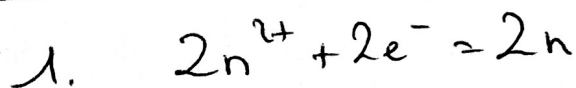


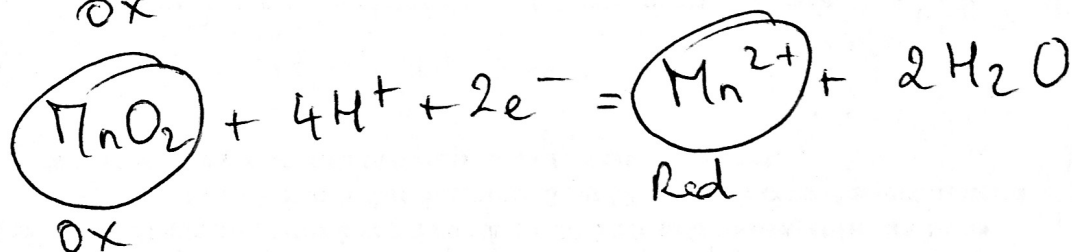
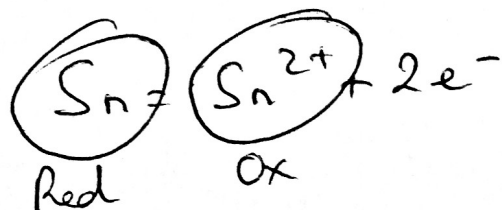
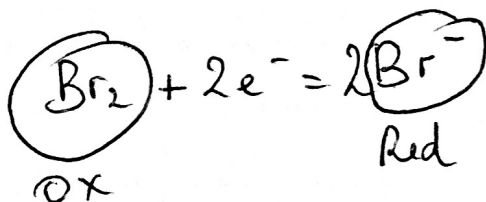
①

Correction Application Equations bilans Ox-red

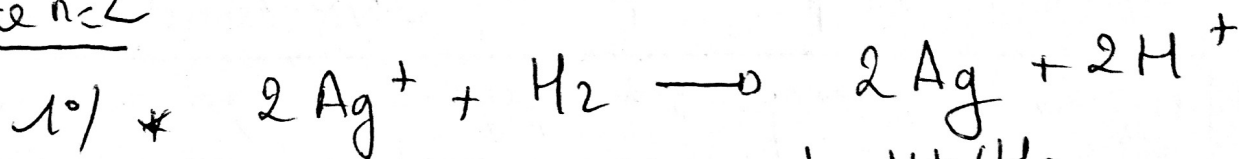
Exercice n°1



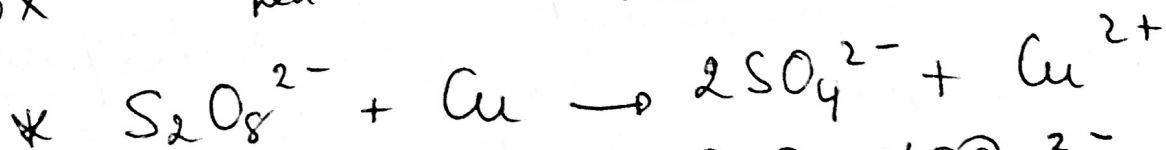
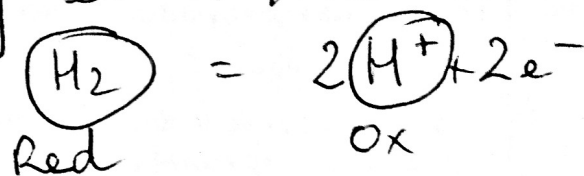
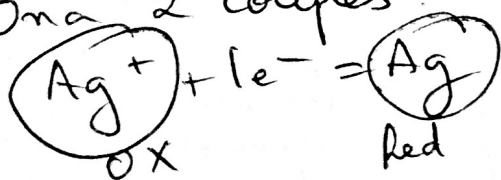
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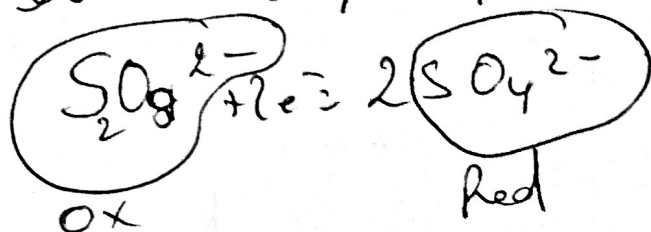
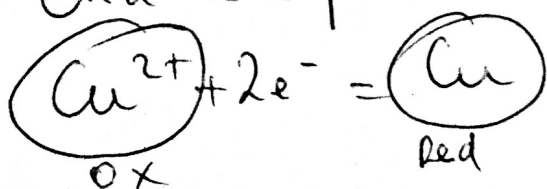
Exercice n°2

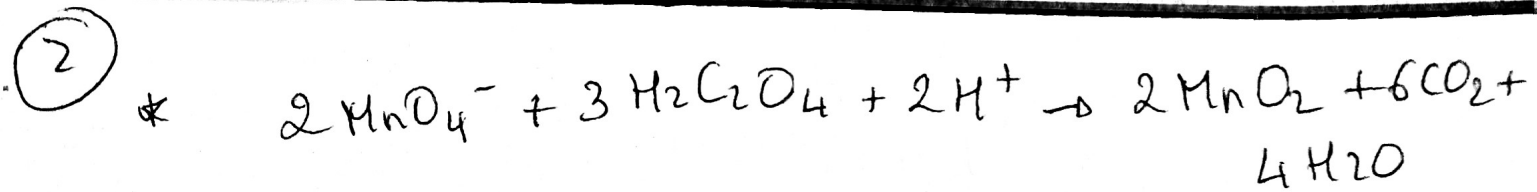


On a 2 couples: Ag^+/Ag et H^+/H_2

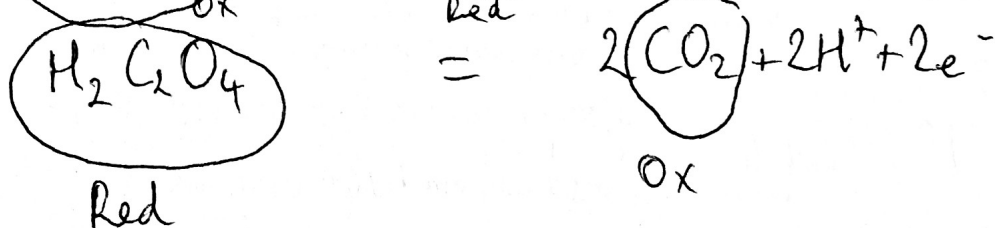
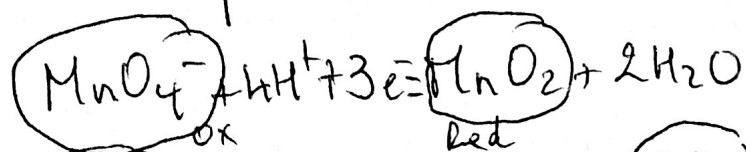


On a 2 couples: Cu^{2+}/Cu et $\text{S}_2\text{O}_8^{2-}/\text{SO}_4^{2-}$





Les couples $\text{MnO}_4^-/\text{MnO}_2$ et $\text{H}_2\text{C}_2\text{O}_4/\text{CO}_2$



2°/ On a le couple I_2/I^-

Les ions I^- sont un REDUCTEUR, ils ne peuvent réagir qu'avec Fe^{2+} lorsqu'il est OXYDANT. I_2/I^- et Fe^{2+}/Fe

I_2 , oxydant peut réagir avec Fe^{2+} lorsqu'il est réducteur I_2/I^- et $\text{Fe}^{3+}/\text{Fe}^{2+}$

3°/ I_2 , oxydant ne peut pas réagir avec Fe^{3+} car il est toujours oxydant

Exercice n°3

1°/ SO_2 réagit avec le diiode I_2
 SO_2 est réducteur et I_2 est oxydant

